

LAMPIRAN B

PERHITUNGAN

1. Perhitungan pembuatan larutan katalis NaOH.

Menghitung penggunaan NaOH 1 M

$$\text{Diketahui : } V = 0,1 \text{ L} = 100 \text{ ml}$$

$$M = 1 \text{ Mol}$$

$$BM = 40 \text{ g/mol}$$

$$\text{Ditanya : Gr} = ?$$

$$\text{Jawab : Gr} = M \times V \times BM$$

$$= 1 \times 100 \times 40$$

$$= 4 \text{ gr}$$

2. Perhitungan pembuatan larutan aktivator NaCl

A. NaCl 0,25 M

$$\text{Diketahui : } V = 0,1 \text{ L} = 100 \text{ ml}$$

$$M = 0,25 \text{ Mol}$$

$$BM = 58,44 \text{ g/mol}$$

$$\text{Ditanya : Gr} = ?$$

$$\text{Jawab : Gr} = M \times V \times BM$$

$$= 0,25 \times 100 \times 58,44$$

$$= 1,46 \text{ gr}$$

B. NaCl 0,5 M

$$\text{Diketahui : } V = 0,1 \text{ L} = 100 \text{ ml}$$

$$M = 0,5 \text{ Mol}$$

$$BM = 58,44 \text{ g/mol}$$

$$\text{Ditanya : Gr} = ?$$

$$\text{Jawab : Gr} = M \times V \times BM$$

$$= 0,5 \times 100 \times 58,44$$

$$= 2,92 \text{ gr}$$

C. NaCl 0,75 M

$$\begin{aligned}
 \text{Diketahui} : V &= 0,1 \text{ L} = 100 \text{ ml} \\
 M &= 0,75 \text{ Mol} \\
 \text{BM} &= 58,44 \text{ g/mol} \\
 \text{Ditanya} : Gr &= ? \\
 \text{Jawab} : Gr &= M \times V \times \text{BM} \\
 &= 0,75 \times 100 \times 58,44 \\
 &= 4,38 \text{ gr}
 \end{aligned}$$

D. NaCl 1 M

$$\begin{aligned}
 \text{Diketahui} : V &= 0,1 \text{ L} = 100 \text{ ml} \\
 M &= 1 \text{ Mol} \\
 \text{BM} &= 58,44 \text{ g/mol} \\
 \text{Ditanya} : Gr &= ? \\
 \text{Jawab} : Gr &= M \times V \times \text{BM} \\
 &= 1 \times 100 \times 58,44 \\
 &= 5,84 \text{ gr}
 \end{aligned}$$

C. Perhitungan % volume gas H₂ menjadi ml

$$\begin{aligned}
 \text{Diketahui} : \% \text{vol} &= 0,88 \\
 \text{Gas Total} &= 700 \text{ ml} \\
 \text{Ditanya} : \text{H}_2 \text{ Total} &= ? \\
 \text{Jawab} : \text{H}_2 \text{ Total} &= \% \text{ volume gas H}_2 \times \text{Gas Total} \\
 &= 0,88 \times 700 \\
 &= 6,16 \text{ ml}
 \end{aligned}$$