

LAMPIRAN B
PERHITUNGAN

1. Penentuan Densitas Sirup Glukosa

a. Penentuan Volume Piknometer

1. Berat Aquadest

$$\begin{aligned} &= (\text{Berat piknometer} + \text{Aquadest}) - (\text{berat piknometer kosong}) \\ &= 56,9189 \text{ gram} - 32,7795 \text{ gram} \\ &= 24,1394 \text{ gram} \end{aligned}$$

2. Volume aquadest = volume piknometer

$$\begin{aligned} &= \frac{\text{Berat aquadest}}{\text{Berat jenis air}} \\ &= \frac{24,1394 \text{ gram}}{1 \text{ gram/ml}} \\ &= 24,1394 \text{ ml} \end{aligned}$$

b. Penentuan Berat Jenis atau Densitas dengan Piknometer

1. Berat Sampel

$$\begin{aligned} &= (\text{Berat piknometer} + \text{Sampel}) - (\text{Berat piknometer kosong}) \\ &= 67,0092 \text{ gram} - 32,7795 \text{ gram} \\ &= 34,2297 \text{ gram} \end{aligned}$$

2. Densitas Sampel, ρ

$$\begin{aligned} \rho &= \frac{\text{Berat Sampel}}{\text{Volume Piknometer}} \\ \rho &= \frac{34,2297 \text{ gram}}{24,1394 \text{ ml}} \\ \rho &= 1,418 \text{ g/ml} \end{aligned}$$

2. Penentuan Kadar Air

- a. Berat sampel (W1) $= 2,5 \text{ gram}$
- b. Berat sampel setelah dikeringkan (W2)
 $= (\text{Berat cawan} + \text{sampel setelah di oven}) - \text{Berat cawan kosong}$
 $= 52,525 \text{ gram} - 50,4 \text{ gram}$
 $= 2,125 \text{ gram}$
- c. Selisih Berat (W3)
 $= W1 - W2$
 $= 2,5 \text{ gram} - 2,125 \text{ gram}$
 $= 0,375 \text{ gram}$
- d. % kadar air

$$= \frac{W3}{W1} \times 100\%$$

$$= \frac{0,375 \text{ gram}}{2,5 \text{ gram}} \times 100\%$$

$$= 15\%$$

3. Penentuan Kadar Abu

- a. Berat Abu
 $= (\text{Berat abu} + \text{krusibel}) - \text{Berat krusibel}$
 $= 30,3140 \text{ gram} - 30,2937 \text{ gram}$
 $= 0,0203 \text{ gram}$
- b. % Abu

$$= \frac{\text{berat abu}}{\text{berat sampel}} \times 100\%$$

$$= \frac{0,0203 \text{ gram}}{2,9 \text{ gram}} \times 100\%$$

$$= 13,6\%$$

4. Penentuan Kadar Glukosa

Dari Kurva Standar Glukosa didapat persamaan berikut :

$$y = 0,0013x + 1,3331$$

Keterangan :

x = Konsentrasi (%)

y = Indeks bias

Maka,

$$x = \frac{y - 1,3331}{0,0013}$$

$$x = \frac{1,4127 - 1,3331}{0,0013}$$

$$x = 61,2308\%$$

5. Pembuatan Larutan HCl 1 N dalam 100 ml

$$M_1 = \frac{\% \times \rho \times 1000}{\text{BM}}$$

$$M_1 = \frac{0,37 \times 1,19 \times 1000}{36,5}$$

$$M_1 = 12,0630 \text{ M}$$

Maka,

$$M_1 \times V_1 = M_2 \times V_2$$

$$V_1 = \frac{M_2 \times V_2}{M_1}$$

$$V_1 = \frac{1 \text{ M} \times 100 \text{ ml}}{12,0630 \text{ M}}$$

$$V_1 = 8,3 \text{ ml}$$

6. Penentuan % Ekstrak dan % Ekstrak Hidrolisis

$$\begin{aligned}\% \text{ Ekstrak} &= \frac{\text{Berat Nira}}{\text{Berat Bengkuang Awal}} \times 100\% \\ &= \frac{1006,5 \text{ g}}{1500 \text{ g}} \times 100\% \\ &= 67,10\%\end{aligned}$$

$$\begin{aligned}\% \text{ Ekstrak Hidrolisis} &= \frac{\text{Berat Nira Hidrolisis}}{\text{Berat Bengkuang Awal}} \times 100\% \\ &= \frac{940,2 \text{ g}}{1500 \text{ g}} \times 100\% \\ &= 62,68\%\end{aligned}$$

7. Penentuan % Yield

$$\begin{aligned}\% \text{ Yield} &= \frac{\text{Berat Produk}}{\text{Berat Bengkuang Awal}} \times 100\% \\ &= \frac{124,8 \text{ g}}{1500 \text{ g}} \times 100\% \\ &= 8,32\%\end{aligned}$$